Scheme of work A

A-level Chemistry 7405

## v1.2

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Scheme of work

## 3.1 Physical Chemistry

This could be taught alongside topics from Organic and/or Inorganic Chemistry. Prior knowledge required from other topics on the specification will be highlighted at the start of each section of the Scheme of Work.

### 3.1.8 Thermodynamics

The further study of thermodynamics builds on the Energetics section and is important in understanding the stability of compounds and why chemical reactions occur. Enthalpy change is linked with entropy change enabling the free-energy change to be calculated.

Prior knowledge:

**AS Chemistry**

- 3.1.4 – Energetics.

3.1.8.1 Born-Haber Cycles

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Definitions of enthalpy changes used in Born–Haber and solution enthalpy cycles.  Using Born–Haber cycles for ionic compounds.  Considering covalent character of ionic compounds.  Using solution enthalpy cycles for ionic compounds. | 1.5 weeks | **Students should be able to:**   * define lattice enthalpy (formation and dissociation), enthalpy of formation, ionisation enthalpy, enthalpy of atomisation, bond enthalpy, electron affinity, enthalpy of solution, hydration enthalpy * draw and use Born–Haber cycles to find missing values of enthalpy changes * comment on the covalent character of an ionic compounds by comparing lattice enthalpies found using Born–Haber cycles with those calculated theoretically using the perfect ionic model. | * Write equations to represent enthalpy changes (AO2 - Apply knowledge and understanding). * Construct Born-Haber cycles and use them to calculate missing enthalpy change values (AO2 - Apply knowledge and understanding; MS2.2 Change the subject of an equation). * Compare and comment on values of enthalpy changes from Born–Haber cycles with those calculated theoretically using the perfect ionic model (AO3 - Analyse, interpret and evaluate data to make judgements). * Construct and use cycles involving the solution of ionic compounds in water to find missing enthalpy change values (AO2 - Apply knowledge and understanding MS2.2 Change the subject of an equation). * Rich question – predict the relative magnitude of the lattice enthalpy of the following compounds: aluminium oxide, potassium oxide, sodium chloride, sodium oxide. * Rich question – for an ionic compound with covalent character, deduce whether the lattice enthalpy will have a greater or smaller magnitude than that calculated theoretically from the perfect ionic model. | * June 2013 Unit 5 Question 1 (QS13.5.01) * June 2013 Unit 5 Question 2 (QS13.5.02) * January 2013 Unit 5 Question 2 (QW13.5.02) * June 2011 Unit 5 Question 1 (QS11.5.01) * January 2010 Unit 5 Question 4 (QW10.5.04) | Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510  Many suitable calculations can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

3.1.8.2 Gibbs free-energy change *ΔG* and entropy change *ΔS*

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| To calculate entropy changes for reactions  To calculate Gibbs free-energy changes and determine whether reactions are feasible at various temperatures | 1.5 weeks | **Students should be able to:**   * describe entropy in terms of disorder * predict whether reactions have an increase or decrease in entropy * calculate the entropy change for a reaction * calculate the Gibbs free-energy change for a reaction at a given temperature * determine whether a reaction is feasible at a given temperature * calculate the temperature at which a reaction becomes feasible * use entropy changes to explain why some endothermic reactions are feasible. | * Rank given substances in terms of entropy (AO2 - Apply knowledge and understanding). * Use entropy values to calculate the entropy change for a reaction (AO2 - Apply knowledge and understanding MS2.2 Change the subject of an equation; MS2.3 Substitute numerical values into algebraic equations). * Predict, where possible, whether reactions have an increase or decrease in entropy (AO2 - Apply knowledge and understanding). * Use the equation *∆G = ∆H – T∆S* to determine whether reactions are feasible at given temperatures, and determine the temperature at which reactions become feasible (AO2 - Apply knowledge and understanding; MS2.2 - Change the subject of an equation; MS2.3 - Substitute numerical values into algebraic equations using appropriate units for physical quantities). * Plot graphs of *∆G* versus T to determine ∆*H* and *∆S* (MS3.3 - Determine the slope and intercept of a linear graph). * Forecast how temperature affects the feasibility of reactions given the sign of the enthalpy and entropy changes (AO2 - Apply knowledge and understanding). * Apply the equation ∆*G* = ∆*H* – *T*∆*S* to state changes to find ∆*H*, ∆*S*, melting and/or boiling points (AO2 - Apply knowledge and understanding; MS2.2 - Change the subject of an equation; MS2.3 - Substitute numerical values into algebraic equations using appropriate units for physical quantities). * Determine ∆*H* and *∆S* for the vaporization of water using a kettle (PS 3.2 - Process and analyse data using appropriate mathematical skills as exemplified in the mathematical appendix for each science). | * June 2013 Unit 5 Question 3 (QS13.5.03) * January 2012 Unit 5 Question 2 (QW12.5.02) * June 2011 Unit 5 Question 2 (QS11.5.02) * June 2010 Unit 5 Question 6 (QS10.5.06) | Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510  RSC Classic Chemical Demonstrations - *∆H* and *∆S* for the vaporization of water using a kettle <http://media.rsc.org/Classic%20Chem%20Demos/CCD-57.pdf>  Many suitable calculations can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

### 3.1.9 Rate Equations

In rate equations, the mathematical relationship between rate of reaction and concentration gives information about the mechanism of a reaction that may occur in several steps

Prior knowledge:

**AS Chemistry**

- 3.1.5 – Kinetics.

3.1.9.1 Rate equations

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand rate equations and order of reaction.  Deduce order of reaction, rate equations and rate constants from rate data.  Describe how the rate constant changes with temperature.  Use the Arrhenius equation. | 0.5 weeks | **Students should be able to:**   * define the terms order of reaction and rate constant * describe how changing concentration of a reagent affects the rate when the order with respect that reagent is 0, 1 or 2 * determine the values and units for rate constants given appropriate data * describe how rate constants change with temperature * perform calculations using the Arrhenius equation * plot straight line graphs of ln k versus 1/T to determine the activation energy of a reaction. | * Describe how changes in concentration will affect reaction rates given the rate equation (AO2 - Apply knowledge and understanding). * Use rate equations to determine reaction rates or rate constants (with units) using initial rate data (AO2 - Apply knowledge and understanding; MS0.0 - Recognise and make use of appropriate units in calculation; MS2.3 – substitute numerical values into algebraic equations; MS2.4 - Solve algebraic equations). * Students use a graph of concentration–time and calculate the rate constant of a zero-order reaction by determination of the gradient. (AO2 - Apply knowledge and understanding; MS3.3 - Determine the slope of a linear graph; MS3.4 - Calculate rate of change from a graph showing a linear relationship). * Students can measure the activation energy for the catalysed and uncatalysed reaction of iodine with peroxodisulphate(VI) ions by experiment and plotting graphs (AO2 - Apply knowledge and understanding; MS3.3 - Determine the slope of a linear graph). | * June 2006 Unit 4 Question 5a and 5b (QS06.4.05) * June 2003 Unit 4 Question 1 (QS03.4.01) | Calculations in AS / A Level Chemistry (Clark) ISBN 9780582411272  *Chemistry Review* article: Establishing a rate equation (Volume 14, edition 2)  Many suitable calculations can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/>  Advanced Practical Chemistry (ILPAC)  ISBN 0719575079 |

3.1.9.2 Determination of rate equation

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand that rate equations have to be determined by experiment.  Link rate equations to mechanisms.  Determine rate using concentration-time graphs.  Use rate-concentration graphs to deduce order for a reagent.  **Required practical 7**  Measure the rate of a reaction by an initial rate method, and a continuous monitoring method. | 2.0 weeks | **Students should be able to:**   * explain that rate equations can only be determined by experiment * use concentration-time graphs to find rates (including initial rates) * use initial rate data to determine rate equations * use rate-concentration data/graphs to find orders of reaction with respect to a reagent * link rate equations to mechanism and determine rate determining steps. | * Determine rate equations, rate constants (with units) using initial rate data (AO2 - Apply knowledge and understanding; MS0.0 - Recognise and make use of appropriate units in calculation; MS2.3 – substitute numerical values into algebraic equations; MS2.4 - Solve algebraic equations). * Students do the iodine clock reaction and determine the order of reaction for a reactant (AO2 - Apply knowledge and understanding; PS 2.4 - Identify variables including those that must be controlled; PS 3.1 - Plot and interpret graphs; PS 3.2 - Process and analyse data using appropriate mathematical skills; MS3.1 - Translate information between graphical, numerical and algebraic forms; MS3.2 - Plot two variables from experimental or other data; MS3.3 - Determine the slope and intercept of a linear graph AT a, k, l). * Students can react calcium carbonate or magnesium with acid of different concentrations and plot volume of gas formed against time for continuous monitoring. Initial rates could be found from these plots and compared (AO2 - Apply knowledge and understanding; PS 2.4 - Identify variables including those that must be controlled; PS 3.1 - Plot and interpret graphs; PS 3.2 - Process and analyse data using appropriate mathematical skills; MS3.1 - Translate information between graphical, numerical and algebraic forms; MS3.2 - Plot two variables from experimental or other data; MS3.3 - Determine the slope and intercept of a linear graph; MS3.4 - Calculate rate of change from a graph showing a linear relationship; MS3.5 - Draw and use the slope of a tangent to a curve as a measure of rate of change; AT a, k, l). * Students can use colorimetry for continuous monitoring experiments (eg bromine + methanoic acid; propanone + iodine) to determine order (AO2 - Apply knowledge and understanding; PS 3.1 - Plot and interpret graphs; PS 3.2 - Process and analyse data using appropriate mathematical skills; MS3.1 - Translate information between graphical, numerical and algebraic forms; MS3.2 - Plot two variables from experimental or other data; MS3.3 - Determine the slope and intercept of a linear graph; MS3.4 - Calculate rate of change from a graph showing a linear relationship; MS3.5 - Draw and use the slope of a tangent to a curve as a measure of rate of change; AT a, k, l). * Students could be given data to plot and interpret in terms of order with respect to a reactant. Alternatively, students could just be given appropriate graphs and asked to derive order(s) (AO2 - Apply knowledge and understanding; MS3.1 - Translate information between graphical, numerical and algebraic forms; MS3.2 - Plot two variables from experimental or other data; MS3.3 - Determine the slope and intercept of a linear graph; MS3.4 - Calculate rate of change from a graph showing a linear relationship; MS3.5 - Draw and use the slope of a tangent to a curve as a measure of rate of change). * Students calculate the rate constant of a zero-order reaction by determining the gradient of a concentration–time graph (MS3.3 - Determine the slope and intercept of a linear graph; MS3.4 - Calculate rate of change from a graph showing a linear relationship). * Students plot concentration–time graphs from collected or supplied data and draw an appropriate best-fit curve. Students draw tangents to such curves to deduce rates at different times (MS3.5 - Draw and use the slope of a tangent to a curve as a measure of rate of change). | * SAM A-level paper 2 (set 1) Q2 * June 2013 Unit 4 Question 1 (QS13.4.01) * January 2013 Unit 4 Question 1 (QW13.4.01) * January 2011 Unit 4 Question 1 (QW11.4.01) * January 2010 Unit 4 Question 3 (QW10.4.03) * January 2006 Unit 4 Question 1 (QW06.4.01) * January 2003 Unit 4 Question 1 (QW03.4.01) | Calculations in AS / A Level Chemistry (Clark) ISBN 9780582411272  *Chemistry Review* article: Establishing a rate equation (Volume 14, edition 2)  ILPAC Unit P5: Chemical Kinetics (free download from [www.nationalstemcentre.org.uk](http://www.nationalstemcentre.org.uk))  Avogadro web site on rate equations: <http://www.avogadro.co.uk/kinetics/rate_equation.htm> |

### 3.1.10 Equilibrium constant *K*pfor homogeneous systems

The further study of equilibria considers how the mathematical expression for the equilibrium constant *K*penables us to calculate how an equilibrium yield will be influenced by the partial pressures of reactants and products. This has important consequences for many industrial processes.

Prior knowledge:

**AS Chemistry**

- 3.1.6 – Chemical equilibria, Le Châtelier’s principle and *K*c

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the concept of and calculate partial pressures using mole fractions.  Write expressions for and calculate *K*pincluding units.  Perform calculations involving *K*p*.*  Predict how changes in conditions affect the position of an equilibrium and the value of *K*p*.*  The effect of a catalyst affects an equilibrium and *K*p*.* | 2.0 weeks | **Students should be able to:**   * calculate equilibrium quantities, mole fractions and partial pressures for equilibrium mixtures * write an expression for *K*p for a reaction and calculate the value of *K*p with units * predict and justify how changes in temperature and pressure affect the position of an equilibrium, and how this may or may not affect the value of *K*p * understand how a catalyst affects an equilibrium and the value of *K*p*.* | * Given initial amounts of substances and one substance at equilibrium, find the quantity of each reagent at equilibrium (AO2 - Apply knowledge and understanding). * Calculate mole fractions and then partial pressures in order to determine *K*p, with units (AO2 - Apply knowledge and understanding; MS2.3 - Substitute numerical values into algebraic equations using appropriate units for physical quantities). * For given equilibria with enthalpy change data, predict the effect on the position of an equilibrium and the value of *K*p (AO2 - Apply knowledge and understanding). | * January 2007 Unit 4 Question 2 (QW04.4.02) * June 2007 Unit 4 Question 1 (QS07.4.01) * January 2008 Unit 4 Question 3 (QW08.4.03) * June 2008 Unit 4 Question 3 (QS08.4.03) * January 2009 Unit 4 Question 3 (QW09.4.03) * June 2009 Unit 4 Question 2 (QS09.4.02) | Calculations for A level Chemistry (Ramsden) ISBN 9780748758395  Many suitable calculations can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

### 3.1.11 Electrode potential and electrochemical cells

Redox reactions take place in electrochemical cells where electrons are transferred from the reducing agent to the oxidising agent indirectly via an external circuit. A potential difference is created that can drive an electric current to do work. Electrochemical cells have very important commercial applications as a portable supply of electricity to power electronic devices such as mobile phones, tablets and laptops. On a larger scale, they can provide energy to power a vehicle.

Prior knowledge:

**AS Chemistry**

* 3.1.7 – Oxidation, reduction and redox equations.

3.1.11.1 Electrode potentials and cells

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| The idea of a cell that has a potential difference being made by combining two half cells (electrodes).  How potentials are measured relative to the Standard Hydrogen Electrode and under standard conditions.  Use the electrochemical series to calculate the EMF of cells and predict the direction of simple redox reactions.  **Required practical 8**  Measuring the EMF of an electrochemical cell. | 1.0 weeks | **Students should be able to:**   * understand that there is a potential difference between two half cells (electrodes) that are joined * use cell notation to represent cells * understand that potentials are measured relative to the Standard Hydrogen Electrode * understand that the potential of an electrode is affected by conditions * know the standard conditions under which potentials are measured * know that electrode potential are listed in order in the electrochemical series * use the electrochemical series to predict the direction of simple redox reactions. | * Students make simple cells and use them to measure EMF and unknown electrode potentials (AO2 - Apply knowledge and understanding; PS 1.1 - Solve problems set in practical contexts; AT j - Set up electrochemical cells and measuring voltages). * Students write the standard cell notation for cells (AO2 - Apply knowledge and understanding). * Students predict how changes in conditions will affect EMF (AO2 - Apply knowledge and understanding). * Students could be asked to plan and carry out an experiment to investigate the effect of changing conditions, such as concentration or temperature, in a voltaic cell such as *Zn|Zn2+||Cu2+| Cu* (AO2 - Apply knowledge and understanding; PS 1.1 - Solve problems set in practical contexts; PS 2.4 - Identify variables including those that must be controlled; AT j). * Students could use *E* values to predict the direction of simple redox reactions, then test these predictions by simple test-tube reactions (AO2 - Apply knowledge and understanding). | * January 2013 Unit 5 Question 7 (QW13.5.07) * January 2012 Unit 5 Question 4 (QW12.5.04) * June 2006 Unit 5 Question 5 (QS06.5.05) * January 2004 Unit 5 Question 4 (QW04.5.04) | Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510  *Chemistry Review* articles: Understanding electrode potentials (Volume 12, edition 1)  Electrode potentials (Volume 15, edition 3)  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

3.1.11.2 Commercial applications of electrochemical cells

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| **Learning objective** | **Time taken** | **Learning Outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| That cells can be used as a source of energy.  That cells can be non-rechargeable or rechargeable.  That fuel cells can be used to generate an electric current.  That there are benefits and risks associated with using these cells. | 1.0 weeks | **Students should be able to:**   * calculate the EMF and cell reaction for a commercial cell given the half-equations * explain how some cells can be recharged * explain how a hydrogen fuel cell works * evaluate the benefits and risks associated with using non-rechargeable, rechargeable and fuel cells. | * Students make simple cells and use them to measure EMF and unknown electrode potentials (AO2 - Apply knowledge and understanding; PS 1.1 - Solve problems set in practical contexts). * Students write the standard cell notation for cells (AO2 - Apply knowledge and understanding). * Students predict how changes in conditions will affect EMF (AO2 - Apply knowledge and understanding). * Students could be asked to plan and carry out an experiment to investigate the effect of changing conditions, such as concentration or temperature, in a voltaic cell such as *Zn|Zn2+||Cu2+| Cu* (AO2 - Apply knowledge and understanding; PS 1.1 - Solve problems set in practical contexts; PS 2.4 - Identify variables including those that must be controlled). * Students could use *E* values to predict the direction of simple redox reactions, then test these predictions by simple test-tube reactions (AO2 - Apply knowledge and understanding). | * June 2013 Unit 5 Question 5 (QS13.5.05) * June 2012 Unit 5 Question 5 (QS12.5.05) * January 2011 Unit 5 Question 5 (QW11.5.05) | Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510  *Chemistry Review* articles: Understanding electrode potentials (Volume 12, edition 1)  Electrode potentials (Volume 15, edition 3)  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

### 3.1.12 Acids and bases

Acids and bases are important in domestic, environmental and industrial contexts. Acidity in aqueous solutions is caused by hydrogen ions and a logarithmic scale, pH, has been devised to measure acidity. Buffer solutions, which can be made from partially neutralised weak acids, resist changes in pH and find many important industrial and biological applications.

Prior knowledge:

**AS Chemistry**

- 3.1.6 – Chemical equilibria, Le Châtelier’s principle and *K*c

3.1.12.1 Brønsted–Lowry acid-base equilibria in aqueous solutions

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| The idea of acids as proton donors and bases as proton acceptors. | 0.2 weeks | **Students should be able to:**   * define Brønsted–Lowry acids and bases * identify species as Brønsted–Lowry acids or bases in proton transfer reactions. | * Identify which species acts as the acid and which as the base in Brønsted-Lowry acid-base reactions (AO2 - Apply knowledge and understanding). | * June 2012 Unit 4 Question 3a and 3b (QS12.4.03) | Theory of acids history websites:  <http://www.bbc.co.uk/dna/ptop/plain/A708257>  <http://pubs.acs.org/subscribe/archive/tcaw/12/i03/pdf/303chronicles.pdf>  RSC acid-base simulator: <http://www.rsc.org/learn-chemistry/resource/res00001457/acid-base-solutions-rsc-funded> |

3.1.12.2 Definition and determination of pH

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Calculate the pH of strong acids from concentration and vice versa. | 0.4 weeks | **Students should be able to:**   * calculate pH of a strong acid from its concentration * calculate the concentration of a strong acid from its pH * calculate the pH of when a strong acid is diluted. | * Identify acids as being strong or weak and monoprotic or diprotic (AO2 - Apply knowledge and understanding). * Calculate the pH of strong acids from the acid concentration, including examples where the acids are diluted (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Calculate the concentration of strong acids from the pH (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). | * June 2009 Unit 4 Question 3a (QS09.4.03) | RSC pH simulator: <http://www.rsc.org/learn-chemistry/resource/res00001458/ph-scale-simulation-rsc-funded>  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |
| Extension |  |  |  |  | Estimate the number of H+ ions in a drop of water  <http://www.rsc.org/learn-chemistry/resource/res00000665/h-ions-in-water> |

3.1.12.3 The ionic product of water *K*w

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Use *K*w to calculate the pH of strong bases. | 0.3 weeks | **Students should be able to:**   * show that *K*w = [H+][OH–] * use *K*w to find the pH of strong bases from its concentration, and vice versa * calculate the pH of water at different temperatures | * Derive the expression *K*w = [H+][OH–] (AO1 - Demonstrate knowledge and understanding). * Calculate the pH of strong bases from the base concentration and vice versa, including dilutions (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Calculate the pH of water at different temperatures (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Explain how the pH and neutrality of water is or is not affected by changes in temperature (AO2 - Apply knowledge and understanding). | * January 2013 Unit 4 Question 2a (QW13.4.02) * June 2011 Unit 4 Question 2a (QS11.4.02) * June 2010 Unit 4 Question 5a and 5b (QS10.4.05) | RSC pH simulator: <http://www.rsc.org/learn-chemistry/resource/res00001458/ph-scale-simulation-rsc-funded>  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

3.1.12.4 Weak acids and bases; *K*a for weak acids

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the term *weak* in relation to acids and bases.  Use *K*a to find the pH of weak acids from the concentration and vice versa.  Relate *Ka* to *pK*a | 0.3 weeks | **Students should be able to:**   * write expressions for *K*afor stated weak acids * perform calculations linking *K*a to concentration and pH * convert *K*avalues to p*K*a and vice versa * calculate the pH of water at different temperatures. | * Explain the difference between strong and weak acids and bases (AO1 - Demonstrate knowledge and understanding). * Derive expressions for *K*a for stated acids (AO1 - Demonstrate knowledge and understanding). * Perform calculations linking *K*a to concentration and pH (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Convert *K*avalues to p*K*a and vice versa, and use these values to rank acids in order of strength (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Measure *K*a of a weak acid by measuring pH at half neutralisation (AO2 - Apply knowledge and understanding; AT c - Measure pH using pH charts, or pH meter, or pH probe on a data logger; PS 4.1 - Know and understand how to use a wide range of experimental and practical instruments, equipment and techniques; AT d). | * January 2012 Unit 4 Question 4b (QW12.4.04) * January 2006 Unit 4 Question 2a and 2b (QW06.4.02) | RSC acid-base simulator: <http://www.rsc.org/learn-chemistry/resource/res00001457/acid-base-solutions-rsc-funded>  RSC pH simulator: <http://www.rsc.org/learn-chemistry/resource/res00001458/ph-scale-simulation-rsc-funded>  Creative problem solving in Chemistry – weak acids: <http://www.rsc.org/learn-chemistry/resource/res00000677/a-weak-acid>  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

3.1.12.5 pH curves, titrations and indicators

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Calculate the pH of the solution formed when strong or weak acids react with strong bases.  Sketch pH curves and choose suitable indicators for titrations.  **Required practical 9**  Investigate how pH changes when a weak acid reacts with a strong base and when a strong acid reacts with a weak base. | 1.0 weeks | **Students should be able to:**   * calculate pH of a mixture of a strong acid with a strong base * calculate the pH of a mixture of a weak acid with a strong base * sketch pH curves for titrations of strong/weak acids with strong/weak bases * choose a suitable indicator for acid-base titrations. | * Perform calculations to find the pH of mixtures of strong/weak acids with strong bases, with either excess acid or base (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Produce pH curves by experiment (AO2 - Apply knowledge and understanding; AT c - Measure pH using pH charts, or pH meter, or pH probe on a data logger; AT d, k, a). * Sketch pH curves for given acid and base combinations, and choose a suitable indicator (AO2 - Apply knowledge and understanding). | * June 2013 Unit 4 Question 3 (QS13.4.03) * June 2011 Unit 4 Question 1 (QS11.4.01) * CHEM4 Specimen Paper Question 3 (QSP 4.03) * June 2005 Unit 4 Question 2 (QS05.4.02) * June 2005 Unit 5 Question 2 (QS05.5.02) * June 2003 Unit 4 Question 3 (QW03.4.03) | RSC pH simulator: <http://www.rsc.org/learn-chemistry/resource/res00001458/ph-scale-simulation-rsc-funded>  pH curve simulators:  <http://chem-ilp.net/labTechniques/AcidBaseIdicatorSimulation.htm>  <http://terpconnect.umd.edu/~toh/models/TitrationDemo.html>  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |
| Extension |  |  | * Write spreadsheets to calculate the pH during a titration and to plot the pH curve (AO2 - Apply knowledge and understanding) |  |  |

3.1.12.6 Buffer action

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know what buffer solutions are, how they are made and what they are used for.  Explain how acidic and basic buffer solutions work.  Calculate the pH of acidic buffer solutions. | 0.6 weeks | **Students should be able to:**   * describe what a buffer solution is and how it is made * explain qualitatively how acidic/basic buffer solutions work * know some uses of buffer solutions * calculate the pH of a buffer solution. | * Describe how buffer solutions are made, how they work and what they are used for (AO2 - Apply knowledge and understanding). * Calculate the pH of a buffer solution given details about quantities of the reagents it is made from, and changes in pH when small amounts of acid/alkali are added to buffer solutions (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude). * Students could prepare a solution of a specific pH and then test the solution to check its pH and buffer action (AO2 - Apply knowledge and understanding; MS0.4 - Use calculators to find and use power, exponential and logarithmic functions; MS2.5 - Use logarithms in relation to quantities that range over several orders of magnitude; AT c - Measure pH using pH charts, or pH meter, or pH probe on a data logger; AT e - Use volumetric flask, including accurate technique for making up a standard solution; PS 1.1 - Solve problems set in practical contexts; PS 4.1 - Know and understand how to use a wide range of experimental and practical instruments, equipment and techniques). | * January 2013 Unit 4 Question 2 (QW13.4.02) * January 2011 Unit 4 Question 2 (QW11.4.02) * CHEM4 Specimen Paper Question 4 (QSP 4.04) * January 2005 Unit 4 Question 8 (QW05.4.08) * January 2002 Unit 4 Question 3 (QW02.4.03) | Sandcastles & mudhuts – buffering action in blood (Hancock) ISBN 9780340543696  Some suitable problems can be found at <http://www.docbrown.info/> and <http://www.chemsheets.co.uk/> (subscription required) |

## 3.2 Inorganic chemistry

### 3.2.4 Properties of Period 3 elements and their oxides

The reactions of the Period 3 elements with oxygen are considered. The pH of the solutions formed when the oxides react with water illustrates further trends in properties across this period. Explanations of these reactions offer opportunities to develop an in-depth understanding of how and why these reactions occur.

Prior knowledge:

**AS Chemistry**

- 3.1.3 – Bonding.

- 3.2.1 – Periodicity.

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Reactions of Na and Mg with water.  Reactions of Na, Mg, Al, Si, P and S with oxygen.  Melting points of period 3 oxides.  Reactions of period 3 oxides with water. | 0.6 weeks | **Students should be able to:**   * describe and write equations for reactions of Na and Mg with water * describe and write equations for reactions of Na, Mg, Al, Si, P and S with oxygen * describe and explain the trend in melting points of period 3 oxides * write equations for the reactions of period 3 oxides with water and describe the pH of the solutions formed * describe the structure and bonding of period 3 oxides, and link this to how they react with water. | * Practical opportunity: react specified period 3 elements with water and oxygen; react specified oxides with water. * Plot a graph of melting points of period 3 oxides and annotate it with explanation of the relative melting points. * Complete tables including equations to show how period 3 elements react with water and/or oxygen, and how period 3 oxides react with water. | * June 2013 Unit 5 Question 4a, 4b and 4c (QS13.5.04) * January 2013 Unit 5 Question 4a, 4b, 4c and 4d (QW13.5.04) * January 2012 Unit 5 Question 3 (QW12.5.03) * January 2011 Unit 5 Question 3 (QW11.5.03) | Youtube video on Period 3 oxides: <https://www.youtube.com/watch?v=D0pNAFjyE6o>  Youtube video of reaction of phosphorus with oxygen: <https://www.youtube.com/watch?v=U6_-EUcswSc&src_vid=mjkuSm__G7s&feature=iv&annotation_id=annotation_323593> |

### 3.2.5 Transition metals

The 3d block contains 10 elements, all of which are metals. Unlike the metals in Groups 1 and 2, the transition metals Ti to Cu form coloured compounds and compounds where the transition metal exists in different oxidation states. Some of these metals are familiar as catalysts. The properties of these elements are studied in this section with opportunities for a wide range of practical investigations.

Prior knowledge:

**AS Chemistry**

- 3.1.1 – Atomic structure (electron structure).

- 3.1.7 – Oxidation, reduction and redox reactions (oxidation states, oxidation, reduction, redox equations).

**3.2.5.1 General properties of the transition metals**

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Give the electron structure of transition metals and their ions.  Know the characteristic properties of transition metals.  Understand the terms complex, ligand co-ordinate bond, and co-ordination number. | 0.2 weeks | **Students should be able to:**   * write the electron structure of first row transition metals and their ions * describe what a transition metal is in terms of electron structure * describe the characteristic properties of transition metals * define the terms ligand, complex, co-ordinate bond and co-ordination number. | * State the electron structure of first row transition metals and their ions (AO1 - Demonstrate knowledge and understanding). * Explain why the elements Ti–Cu have properties characteristic of transition metals, and what those characteristics are (AO1 - Demonstrate knowledge and understanding). * Identify the oxidation state of the metal, the ligands and co-ordination number in a series of complexes (AO2 - Apply knowledge and understanding). * Identify an element from the series Ti–Cu and find examples for that element to confirm its characteristic properties (AO3 - Analyse, interpret and evaluate scientific information, ideas and evidence). | * January 2005 Unit 5 Question 6a (QW05.5.06) * June 2010 Unit 5 Question 4a and 4b (QS10.5.06) * January 2011 Unit 5 Question 4a and 4b (QW11.5.04) | *Chemistry Review* article: Vanadium (Volume 19, edition 4) |

3.2.5.2 Substitution reactions

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the different types of ligands.  Understand ligand exchange.  Know about oxygen transfer by haemoglobin.  Understand the chelate effect. | 1.0 weeks | **Students should be able to:**   * explain the difference between, and give examples of monodentate, bidentate and multidentate ligands * explain what happens in a ligand substitution (exchange) reaction and why there may be a change in co-ordination number * describe what haem is, how oxygen is carried in blood and why carbon monoxide is toxic * describe and explain the chelate effect in terms of enthalpy and entropy changes. | * Give examples of monodentate, bidentate and multidentate ligands (AO1 - Demonstrate knowledge and understanding). * Students should carry out substitution reactions of metal aqua complexes with monodentate ligands (from ammonia and concentrated hydrochloric acid) to consider whether there is a change in co-ordination number and whether all the water ligands are substituted (AO2 - Apply knowledge and understanding; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Students could carry out test-tube reactions of complexes with monodentate, bidentate and multidentate ligands to compare ease of substitution (AO2 - Apply knowledge and understanding; PS 1.1 - Solve problems set in practical contexts; AT d, k). | * June 2010 Unit 5 Question 4a, 4b, 4d and 4e (QS10.5.06) * January 2005 Unit 5 Question 6b (QW05.5.06) * June 2004 Unit 5 Question 4b (QS04.5.04) * June 2002 Unit 5 Question 6 (QS02.5.06) | Molecule of month article on EDTA [http://www.chm.bris.ac.uk/motm/edta/edtah.htm - Practical Uses of EDTA](http://www.chm.bris.ac.uk/motm/edta/edtah.htm#Practical Uses of EDTA)  RSC article on uses of EDTA  <http://www.rsc.org/chemistryworld/podcast/CIIEcompounds/transcripts/EDTA.asp> |

3.2.5.3 Shapes of complex ions

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know the shapes of complexes with 2/4/6 ligands.  Understand how complexes can show *cis–trans (E–Z)* or optical isomerism. | 0.3 weeks | **Students should be able to:**   * sketch examples of octahedral, tetrahedral, square planar and linear complexes * know how some complexes can show *cis-trans (E–Z)­* or optical isomerism * know the complexes in cisplatin and Tollen’s reagent. | * Give examples of and sketch the shape of octahedral, tetrahedral, square planar and linear complexes (AO1 - Demonstrate knowledge and understanding; MS4.1 - Use angles and shapes in regular 2D and 3D structures; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). * Explain how *cis–trans (E–Z)* isomerism arises in some octahedral and square planar complexes, including cisplatin, and draw the isomers (AO2 - Apply knowledge and understanding). * Explain how opticalisomerism arises in some octahedral complexes with bidentate ligands, and draw the isomers (AO2 - Apply knowledge and understanding). * Students can use Molymod kits to make models of isomers (AO2 - Apply knowledge and understanding; MS4.1 - Use angles and shapes in regular 2D and 3D structures; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). | * January 2011 Unit 5 Question 4a, 4b and 4c (QW11.5.04) * January 2004 Unit 5 Question 10b (QW04.5.10) * June 2003 Unit 5 Question 3 (QS03.5.03) | Molymod molecular models  Shapes viewer (including inorganic complexes)  [https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/#](https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/) |

3.2.5.4 Formation of coloured ions

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand why transition metal ions are coloured and what affects the colour.  Use colorimetry to measure concentration of solutions. | 0.6 weeks | **Students should be able to:**   * explain why transition metal complexes are coloured * describe factors that affect the colour of transition metal ions * describe how colorimetry can be used to find the concentration of coloured ions in solution. | * Explain using diagrams and the equation *∆E* *= hν (= hc/λ)* why transition metal complexes are coloured and what factors affect the colour (AO1 - Demonstrate knowledge and understanding). * Use a graph of absorption versus concentration to determine the concentration of the solution. * Use a colorimeter to produce a calibration curve and then find the concentration of a coloured solution, eg containing copper(II) ions (AO2 - Apply knowledge and understanding; AT a - Use appropriate apparatus to record a range of measurements; AT d - Use laboratory apparatus for a variety of experimental techniques; PS 3.1 - Plot and interpret graphs; MS3.2 - Plot two variables from experimental or other data). | * June 2013 Unit 5 Question 6 (QS13.5.06) * January 2012 Unit 5 Question 7a and 7b (QW12.5.07) * Specimen Paper Unit 5 Question 6 (QSP.5.6) * June 2002 Unit 5 Question 3 (QW02.5.03) | Colorimetric determination of a copper ore:  <http://www.nuffieldfoundation.org/practical-chemistry/colorimetric-determination-copper-ore>  *Chemistry Review* article: Colorimetry (Volume 12, edition 3)  RSC booklet on colorimetry from Gifted & Talented Chemistry: <http://www.rsc.org/learn-chemistry/resource/res00000847/spectroscopy>  RSC Spectral School with range of resources: <http://www.rsc.org/learn-chemistry/collections/spectroscopy> |

3.2.5.5 Variable oxidation states

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know what happens when vanadate(V) is reduced by zinc in acidic solution.  How the redox potential for a transition metal is affected by the pH and ligand.  The reduction of silver (I) in Tollen’s reagent to test for aldehydes.  Redox titrations, including calculations, of MnO4– with Fe2+ and C2O42– in acidic solution. | 2.5 weeks | **Students should be able to:**   * describe and explain what happens when vanadate(V) ions are reduced by zinc in acidic solution * understand how the redox potential of a transition metal ion is affected by changes in pH and ligand * describe and explain the use of Ag(NH3)2+ in Tollen’s reagent to distinguish between aldehydes and ketones * perform titrations and associated calculations for redox reactions of MnO4– with Fe2+ and C2O42–in acidic solution. | * Students should react an acidified solution of ammonium vanadate(V) with zinc to observe colour changes , identify vanadium species and write redox reactions for each reduction reaction (AO1 - Demonstrate knowledge and understanding; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Compare redox potentials for Cr3+ at different pH values and different ligands (AO2 - Apply knowledge and understanding). * Test aldehydes and ketones with Tollens reagent (AO1 - Demonstrate knowledge and understanding; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Carry out redox titrations, including associated titrations, of Fe2+ with MnO4– in acidic solution (eg analysis of iron in iron tablets, analysis of iron in lawn sand, analysis of iron in steel, finding the *Mr* of hydrated ammonium (II) sulfate) (AO2 - Apply knowledge and understanding; AT d - Use laboratory apparatus for a variety of experimental techniques including titration, using burette and pipette; AT e - Use volumetric flask, including accurate technique for making up a standard solution; AT k; PS 2.3 - Evaluate results and draw conclusions with reference to measurement uncertainties and errors; MS2.3 – substitute numerical values into algebraic equations). * Carry out redox titrations, including associated titrations, of C2O42– with MnO4– in acidic solution (eg finding the *Mr* of ethanedioic acid) (AO2 - Apply knowledge and understanding; AT d - Use laboratory apparatus for a variety of experimental techniques including titration, using burette and pipette; AT e - Use volumetric flask, including accurate technique for making up a standard solution; AT k; PS 2.3 - Evaluate results and draw conclusions with reference to measurement uncertainties and errors; MS2.3 – substitute numerical values into algebraic equations). | * June 2006 Unit 5 Question 1 (QW06.5.01) * June 2005 Unit 5 Question 5 (QS05.5.05) * June 2003 Unit 5 Question 2 (QS03.5.02) * January 2003 Unit 5 Question 7b (QW03.5.07) | Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510  ILPAC Advanced Practical Chemistry ISBN 9780719575075 |

3.2.5.6 Catalysts

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand what heterogeneous catalysts are and how they work, including examples and how they can become poisoned.  Understand what homogeneous catalysts are, with specific examples. | 1.0 weeks | **Students should be able to:**   * describe what a heterogeneous catalyst is and the role of active sites and the support medium * explain, with the aid of equations, how V2O5, acts as a catalyst in the Contact Process * describe the use of Fe in the Haber process * explain how heterogeneous catalysts can become poisoned * describe what a homogeneous catalyst is and how reactions proceed through an intermediate species * describe, with the aid of equations, how Fe2+ catalyses the reaction between I– and S2O82– * describe, with the aid of equations, how Mn2+ catalyses the reaction between C2O42– and MnO4– | * Create a set of notes on how heterogeneous and homogeneous catalysts work, including the specific examples required: * V2O5 in the contact process. * Fe2+ ions in the reaction of I– with S2O82–. * Mn2+ ions in the reaction of C2O42– and MnO4– (AO1 - Demonstrate knowledge and understanding). * Students could investigate Mn2+ as the autocatalyst in the reaction between ethanedioic acid and acidified potassium manganate(VII) (AO2 - Apply knowledge and understanding). | * January 2013 Unit 5 Question 6 (QW13.5.06) * January 2012 Unit 5 Question 6 (QW12.5.06) * January 2011 Unit 5 Question 4d (QW11.5.04) * January 2010 Unit 5 Question 1 (QW10.5.01) * June 2006 Unit 5 Question 9 (QS06.5.09) | *Chemistry Review* article: Catalysts: getting chemistry going (Volume 20, edition 3)  *Chemistry Review* article: Catalysts: heterogeneous catalysis (Volume 23, edition 1)  *Chemistry Review* article: Catalysts: homogeneous catalysis (Volume 23, edition 3) |

### 3.2.6 Reactions of ions in aqueous solution

The reactions of transition metal ions in aqueous solution provide a practical opportunity for students to show and to understand how transition metal ions can be identified by test-tube reactions in the laboratory.

Prior knowledge:

**AS Chemistry**

- 3.1.7 – Oxidation, reduction and redox reactions (oxidation states, oxidation, reduction, redox equations).

**A-level Chemistry**

- 3.2.5 – Transition metals.

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| The nature of meta-aqua ions.  The relative acidity of metal-aqua ions.  The reaction of metal-aqua ions (Fe2+, Cu2+, Al3+, Fe3+) with bases OH–, NH3, CO32–.  The character of metal hydroxides as basic or amphoteric.  **Required practical 11**  Carry out simple test-tube reactions to identify transition metal ions in aqueous solution. | 1.0 weeks | **Students should be able to:**   * understand that metal ions exist as metal-aqua ions in aqueous solution * the hydrolysis of metal-aqua ions in aqueous solution giving acidic solutions * explain why [M(H2O)6]3+ ions are more acidic than [M(H2O)6]2+ions * describe and explain reactions of [M(H2O)6]2+ (M = Cu, Fe) and [M(H2O)6]3+ (M = Al, Fe) with the bases OH–, NH3, CO32– * describe if and how metal hydroxides (Cu(II), Fe(II), Al(III), Fe(III)) react with H+ and OH–, and so whether these metal hydroxides are basic or amphoteric. | * Students could measure the pH of solution of metal aqua ions (of equal concentration) and explain the difference in pH (AO2 - Apply knowledge and understanding; AT c,d,k). * Students could complete a series of test-tube reactions of iron(II) and iron(III) ions with reagents such as Mg, Na2CO3 to exemplify the difference in pH; AT d,k). * Students could carry out test-tube reactions of metal-aqua ions with NaOH, NH3 and Na2CO3 (AO2 - Apply knowledge and understanding; AT d,k). * Students could identify unknown substances (containing cations and anions on the specification) using reagents (AO2 - Apply knowledge and understanding; AT d,k). * Students could produce precipitates of metal hydroxides and then test how they react with acid and alkali to determine whether they are basic or amphoteric (AO2 - Apply knowledge and understanding; AT d,k). | * January 2013 Unit 5 Question 8 (QW13.5.08) * Specimen Paper CHM5 Question 8 (QSP.5.08) * June 2004 Unit 5 Question 4 (QW04.5.04) | Complexes and First Row Transition Metals (Nicholls) ISBN 9780333170885  AQA Reactions of metal ions in aqueous solution resource:  <http://filestore.aqa.org.uk/resources/chemistry/AQA-7405-REACTIONS-OF-METAL-IONS.PDF> |

## 3.3 Organic Chemistry

### 3.3.7 Optical isomerism

Compounds that contain an asymmetric carbon atom form stereoisomers that differ in their effect on plane polarised light. This type of isomerism is called optical isomerism.

Prior knowledge:

**AS Chemistry**

- 3.3.1.3 – Isomerism.

**A-level Chemistry**

- 3.3.8 – Aldehydes and ketones (the best example of how a racemic mixture forms comes from the reaction of many aldehydes and ketones with HCN; two alternative strategies are (a) teach 3.3.7 first but teach the point about formation of racemic mixtures during 3.3.8, or (b) teach 3.3.8 before 3.3.7).

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| The cause and nature of optical isomerism.  The similarities and differences in the properties of enantiomers.  The formation of racemic mixtures. | 0.4 weeks | **Students should be able to:**   * explain the cause of optical isomerism * identify molecules that exhibit optical isomerism/that are optically active. * draw pairs of optical isomers in 3D * describe how enantiomers affect plane polarised light * explain what a racemic mixture is, how they can be formed, and their effect on plane polarised light. | * Students make models of mirror image molecules of some chiral and non-chiral molecules to see if they are non-superimposable or not (AO2 - Apply knowledge and understanding; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). * Students identify whether molecules exhibit optical isomerism, and where they do draw the two enantiomers in 3D (AO2 - Apply knowledge and understanding; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). * Students could see how passing polarised light through a solution of sucrose affects the plane of the light (PS 1.2 - Apply scientific knowledge to practical contexts). * Students could use Molymod models to show how a racemic mixture is formed when ethanal reacts with HCN (AO2 - Apply knowledge and understanding; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). | * January 2005 Unit 4 Question 3d (QW05.4.03) * June 2002 Unit 4 Question 5 (QW02.4.05) | Molymod models  *Chemistry Review* article: Looking in the mirror (Volume 10, edition 3) |

### 3.3.8 Aldehydes and ketones

Aldehydes, ketones, carboxylic acids and their derivatives all contain the carbonyl group which is attacked by nucleophiles. This section includes the addition reactions of aldehydes and ketones

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature

- 3.3.1.2 – Reaction mechanisms

- 3.3.5.2 – Oxidation of alcohols

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know about the oxidation of aldehydes.  Know about the reduction of aldehydes and ketones with NaBH4, including mechanism.  Know about the reaction of aldehydes and ketones with KCN then acid, including mechanism. | 0.6 weeks | **Students should be able to:**   * write equations and know reagents and conditions to oxidise aldehydes to carboxylic acids * know how to distinguish aldehydes and ketones * write equations, know reagents and conditions and outline the mechanism to reduce aldehydes and ketones to alcohols with NaBH4 * write equations, know reagents and conditions and outline the mechanism for reaction of aldehydes and ketones with KCN and acid * understand why reaction of aldehydes and ketones with KCN followed by acid can form a racemic mixture * students understand the hazards of using KCN | * Students write equations for the oxidation of aldehydes (using reagents acidified potassium dichromate(VI) / Tollen’s reagent / Fehling’s solution) (AO2 - Apply knowledge and understanding). * Students could carry out test-tube reactions of Tollens’ reagent and Fehling’s solution to distinguish aldehydes and ketones (AO2 - Apply knowledge and understanding; AT b - Use water bath for heating; AT d - Use laboratory apparatus for a variety of experimental techniques including qualitative tests organic functional groups; AT k). * Students write equations and mechanisms for the reduction of aldehydes and ketones using NaBH4 (AO2 - Apply knowledge and understanding). * Students write equations and mechanisms for the reaction of aldehydes and ketones with KCN followed by acid (AO2 - Apply knowledge and understanding). * Students could use Molymod models to show how a racemic mixture is formed when ethanol reacts with HCN (AO2 - Apply knowledge and understanding; MS4.2 - Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects; MS4.3 - Understand the symmetry of 2D and 3D shapes). * Students could research why KCN/HCN are highly toxic (AO3 - Analyse, interpret and evaluate scientific information). | * January 2010 Unit 4 Question 4 (QW10.4.04) * June 2005 Unit 4 Question 3a (QS05.4.03) * June 2004 Unit 4 Question 6d and 6e (QS04.4.06) * January 2002 Unit 4 Question 6a (QW02.4.06) | Molymod models  Giant silver mirror <http://www.nuffieldfoundation.org/practical-chemistry/giant-silver-mirror>  RSC mechanisms resource: <http://www.rsc.org/learn-chemistry/resource/res00000638/curly-arrows-and-stereoselectivity-in-organic-reactions>  Mechanism animations <http://science.jbpub.com/organic/movies/> |

### 3.3.9 Carboxylic acids and derivatives

Carboxylic acids are weak acids but strong enough to liberate carbon dioxide from carbonates. Esters occur naturally in vegetable oils and animal fats. Important products obtained from esters include biodiesel, soap and glycerol.

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature.

- 3.3.1.2 – Reaction mechanisms.

- 3.3.5.2 – Oxidation of alcohols.

3.3.9.1 Carboxylic acids and esters

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Draw the structure of and name carboxylic acids and esters.  Know that carboxylic acids are weak acids.  Know how esters are made from carboxylic acids.  Know some uses of esters.  Know how esters are hydrolysed.  Know that vegetable oils and animal fats are esters of fatty acids and glycerol  Know how soap and biodiesel are made from vegetable oil and animals fats | 1.5 weeks | **Students should be able to:**   * draw the structure of and name carboxylic acids and esters * know how carboxylic acids react with carbonates * write equations for the reaction of carboxylic acids with alcohols to form esters * know some common uses of esters * write equations for the hydrolysis of esters in acidic or alkaline conditions * understand the structure of animals fats and vegetable oils * know how soap and biodiesel are made and write equations for these reactions for specified fats/oils. | * Students draw and name carboxylic acids and esters (AO2 - Apply knowledge and understanding). * Students write equations for, and make esters by reactions of alcohols with carboxylic acids in test tubes; or an ester could be collected and purified using a separating funnel and distillation (AO2 - Apply knowledge and understanding; AT g - Purify a liquid product, including use of separating funnel; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Students research uses of esters and the presence of esters in fruit (AO2 - Apply knowledge and understanding). * Students write equations for the hydrolysis of given esters in acidic and alkaline conditions (AO2 - Apply knowledge and understanding). * Students make soap by hydrolysis of castor oil (AO2 - Apply knowledge and understanding; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Students make biodiesel (AO2 - Apply knowledge and understanding; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Students write equations for production of soap and/or biodiesel from specified fats/oils (AO2 - Apply knowledge and understanding). * Students could identify an unknown ester by determination of boiling point followed by hydrolysis and then purifying and finding the melting point of the carboxylic acid formed (eg for example methyl benzoate) (AO3 - Analyse, interpret and evaluate scientific information; AT d - Use laboratory apparatus for a variety of experimental techniques including distillation and heating under reflux, including setting up glassware using retort stand and clamps; AT d - Use laboratory apparatus for a variety of experimental techniques including filtration, including use of fluted filter paper, or filtration under reduced pressure; AT k). | * January 2013 Unit 4 Question 3a (QW13.4.03) * June 2010 Unit 4 Question 7a and 7d (QS10.4.07) * January 2010 Unit 4 Question 5 (QW10.4.05) * June 2005 Unit 1 Question 1a and 1d (QS05.4.01) | Making soap from castor oil:  <http://www.nuffieldfoundation.org/practical-chemistry/making-soaps-and-detergents>  Method and guidance for making biodiesel – CLEAPSS leaflet PS 67-10  Molecule of the month: Esters in fruits <http://www.chm.bris.ac.uk/motm/ethylacetate/ethylv.htm>  Biofuels website: <http://www.thesolarspark.co.uk/the-science/renewable-energy/bio/>  Biofuels website:  <http://www.biofuels.co.uk/>  Press report about problems with biofuels: <http://www.telegraph.co.uk/earth/energy/biofuels/10520736/The-great-biofuels-scandal.html>  BP biofuels resources: <http://bpes.bp.com/secondary-resources/science/ages-14-to-16/energy-electricity-and-forces/biofuels-and-the-future> |

3.3.9.2 Acylation

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Draw the structure of and name acid anhydrides, acyl chlorides and amides.  Understand acylation reactions of water, alcohols, ammonia and amines with acyl chlorides and acid anhydrides, including the mechanism for acyl chlorides.  **Required practical 10**  Preparation of  - a pure organic solid and test of its purity  - a pure organic liquid. | 2.0 weeks | **Students should be able to:**   * draw the structure of and name acid anhydrides, acyl chlorides and amides * identify the products of and write equations for acylation reactions of water, alcohols, ammonia and amines with acyl chlorides and acid anhydrides * outline the mechanism for the acylation reactions of acyl chlorides * state advantages of using ethanoic anhydride rather than ethanoyl chloride in the production of aspirin * prepare and purity an organic solid and test its purity. | * Students draw and name acid anhydrides, acyl chlorides and amides (AO2 - Apply knowledge and understanding). * Students write equations and outline mechanisms for acylation reactions of water, alcohols, ammonia and amines with acyl chlorides and acid anhydrides; some of these reactions could be demonstrated. * Students prepare, purify and test the purity of aspirin by melting point determination (AO2 - Apply knowledge and understanding; AT d - Use laboratory apparatus for a variety of experimental techniques including distillation and heating under reflux, including setting up glassware using retort stand and clamps; AT d - Use laboratory apparatus for a variety of experimental techniques including filtration, including use of fluted filter paper, or filtration under reduced pressure; AT k. | * January 2012 Unit 4 Question 10a (QW12.4.10) * June 2006 Unit 4 Question 1 (QS06.4.01) * June 2005 Unit 4 Question 7 (QS05.4.07) * June 2003 Unit 5 Question 8b (QS03.5.08) | RSC resource on aspirin: <http://www.rsc.org/learn-chemistry/resource/res00000056/aspirin>  Aspirin Pre-lab Screen Experiment:  http://www.rsc.org/learn-chemistry/resource/res00001644/aspirin-screen-experiment  RSC mechanisms resource: <http://www.rsc.org/learn-chemistry/resource/res00000638/curly-arrows-and-stereoselectivity-in-organic-reactions>  Mechanism animations <http://science.jbpub.com/organic/movies/> |

### 3.3.10 Aromatic Chemistry

Aromatic chemistry takes benzene as an example of this type of molecule and looks at the structure of the benzene ring and its substitution reactions

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature.

- 3.3.1.2 – Reaction mechanisms.

3.3.10.1 Bonding

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the structure of benzene and evidence for delocalisation. | 0.2 weeks | **Students should be able to:**   * describe the structure of benzene and explain how delocalisation makes benzene more stable than the theoretical cyclohexa-1,3,5-triene * use thermochemical evidence from enthalpies of hydrogenation to account for this extra stability * explain why benzene undergoes substitution reactions in preference to addition reactions. | * Name a range of aromatic compounds with common functional groups (AO2 - Apply knowledge and understanding). * Draw enthalpy diagrams to show the relative stability of cyclohexane, cyclohexene, cyclohexa-1,4-diene, benzene and the theoretical cyclohexa-1,3,5-triene (AO2 - Apply knowledge and understanding). | * June 2011 Unit 4 Question 8a (QS11.4.08) * January 2004 Unit 4 Question 7a (QW04.4.07) | *Chemistry Review* article: The structure of benzene (Volume 1, edition 1)  *Chemistry Review* article: Who discovered the structure of benzene (Volume 5, edition 1) |

3.3.10.2 Electrophilic substitution

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know nitration and Friedel-Crafts acylation reactions of aromatic compounds, including the mechanism and usefulness. | 1.0 weeks | **Students should be able to:**   * write equations and outline mechanisms for nitration and Friedel-Crafts acylation reactions of aromatic compounds. (including equations for the formation of electrophiles) * understand the usefulness of nitration and Friedel-Crafts acylation reactions | * Write equations (including for the formation of electrophiles) and mechanisms for nitration and Friedel-Crafts acylation reactions given the starting material and products (AO2 - Apply knowledge and understanding). * Students could carry out the preparation of methyl 3-nitrobenzoate by nitration of methyl benzoate, purification by recrystallisation and determination of melting point (AT d - Use laboratory apparatus for a variety of experimental techniques including filtration, including use of fluted filter paper, or filtration under reduced pressure; AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances). | * January 2012 Unit 4 Question 9a (QW12.4.09) * January 2011 Unit 4 Question 6 (QW11.4.06) * June 2010 Unit 4 Question 8 b) (QS10.4.08) * January 2006 Unit 4 Question 7 (QW06.4.07) * June 2011 Unit 4 Question 8b | *Chemistry review* article: Probably the most important reactions in the world (Volume 15, edition 2) |

### 3.3.11 Amines

Amines are compounds based on ammonia where hydrogen atoms have been replaced by alkyl or aryl groups. This section includes their reactions as nucleophiles

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature.

- 3.3.1.2 – Reaction mechanisms.

- 3.3.3.1 – Nucleophilic substitution.

3.3.11.1 Preparation

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Know how primary aliphatic amines are made from halogenoalkanes and from nitriles.  Know how aromatic amines are produced and their use in making dyes. | 0.2 weeks | **Students should be able to:**   * write equations and give conditions for the preparation of primary aliphatic amines from both halogenoalkanes and nitriles * write equations and give conditions for the production of aromatic amines and identify their use in making dyes. | * Identify reagents and conditions and write equations to make specified primary aliphatic amines from halogenoalkanes and nitriles (AO2 - Apply knowledge and understanding). * Identify reagents and conditions and write equations to make specified aromatic amines (AO2 - Apply knowledge and understanding). * Research the use of aromatic amines in making dyes (AO3 - Analyse, interpret and evaluate scientific information). | * June 2013 Unit 4 Question 8 (QS13.4.08) * June 2005 Unit 4 Question 5b (QS05.4.05) * January 2005 Unit 4 Question 1 (QW05.4.01) * June 2004 Unit 4 Question 4a and 4b (QS04.4.04) * January 2004 Unit 4 Question 8 (QW04.4.08) | *Chemistry Review* article: Get real: chemistry in fashion (Volume 11, edition 3) |

3.3.11.2 Base properties

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Compare the base strength of amines. | 0.2 weeks | **Students should be able to:**   * place amines in order of base strength and explain this order. | * Given pairs of amines, students should identify the stronger base giving reasons for their choice (AO2 - Apply knowledge and understanding). | * January 2005 Unit 4 Question 1d (QW05.4.01) * June 2004 Unit 4 Question 4c (QS04.4.04) * January 2003 Unit 4 Question 6 (QW03.4.06) * June 2013 Unit 4 Question 9a | Data books with base strength values:  Nuffield Science Data Book (free download):  <http://www.nationalstemcentre.org.uk/elibrary/resource/3402/nuffield-advanced-science-book-of-data-second-edition>  Chemistry Data Book (Starck, Wallace, McGlashan) ISBN: 9780719539510 |

3.3.11.3 Nucleophilic properties

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand how amines react with halogenoalkanes, acyl chlorides and acid anhydrides, including mechanisms. | 0.5 weeks | **Students should be able to:**   * identify the various amines and quaternary ammonium salts formed when ammonia and amines react with halogenoalkanes * give the mechanism for reactions of ammonia and amines with halogenoalkanes * recognise the use of quaternary ammonium salts * identify the products of and write equations for acylation reactions of ammonia and amines with acyl chlorides and acid anhydrides * outline the mechanism for the acylation reactions | * Identify the amines and quaternary ammonium salts that can be formed when ammonia and amines react with halogenoalkanes and how changing conditions can affect the main product; outline the mechanism to form these products (AO2 - Apply knowledge and understanding). * Students could research the use of quaternary ammonium salts (AO3 - Analyse, interpret and evaluate scientific information). * Students write equations and mechanisms for acylation reactions of water, alcohols, ammonia and amines with acyl chlorides and acid anhydrides; some of these reactions could be demonstrated (AO2 - Apply knowledge and understanding). * Practical opportunity: The preparation of N-phenylethanamide. | * January 2006 Unit 4 Question 5 (QW06.4.05) * January 2004 Unit 4 Question 8 (QW04.4.08) * January 2003 Unit 4 Question 6 (QW06.4.05) | *Chemistry Review* article: Two in one: the chemistry of shampoo and conditioner (Volume 22, edition 3) |

### 3.3.12 Polymers

The study of polymers is extended to include condensation polymers. The ways in which condensation polymers are formed are studied, together with their properties and typical uses. Problems associated with the reuse or disposal of both addition and condensation polymers are considered

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature.

- 3.3.4.3 – Addition polymers.

3.3.12.1 Condensation polymers

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand how condensation polymers are formed including linkages in polyesters and polyamides.  Identify the repeating unit given monomer(s) and vice versa. | 0.5 weeks | **Students should be able to:**   * identify the repeating unit and linkages in polyesters and polyamides given the monomer(s) * identify monomer(s) needed the make a condensation polymer given the repeating unit * know the repeating units in Terylene, nylon 6,6 and Kevlar * know some uses of condensation polymers * explain the nature of the intermolecular forces between molecules of condensation polymers. | * Draw the structure of repeating units in polyesters and polyamides given the monomer(s) and vice versa (AO2 - Apply knowledge and understanding). * Students could make nylon 6,6 (AO2 - Apply knowledge and understanding); AT k - Safely and carefully handle solids and liquids, including corrosive, irritant, flammable and toxic substances; AT d). * Students could each make a model of a monomer using Molymods and then students collectively join them together to make a long polymer chain (AO2 - Apply knowledge and understanding). | * January 2012 Unit 4 Question 8b (QW12.4.08) * June 2011 Unit 4 Question 4a (QS11.4.04) * June 2006 Unit 4 Question 4a (QS06.4.04) * June 2004 Unit 4 Question 5 (QS04.4.05) * June 2003 Unit 4 Question 5b) (QS03.4.05) | Molymods  RSC resource on nylon: <http://www.rsc.org/learn-chemistry/resource/res00000026/nylon>  The discovery of Nylon  <http://www.rsc.org/learn-chemistry/resource/res00000034/anecdotes-nylon>  Making nylon:  <http://www.rsc.org/learn-chemistry/resource/res00000755/making-nylon-the-nylon-rope-trick>  Sandcastles and mudhuts section 27 – Spare Parts (Hancock) ISBN 9780340543696  *Chemistry Review* article: Tougher than a speeding bullet (Volume 13, edition 4)  *Chemistry Review* article: Polyesters: plastics of the future (Volume 17, edition 1)  *Chemistry Review* article: Kevlar and composites (Volume 20, edition 2)  *Chemistry Review* article: Kevlar – miracle material (Volume 22, edition 4) |

3.3.12.2 Biodegradability and disposal of polymers

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand why polyalkenes are not biodegradable.  Understand why polyesters and polyamides are biodegradable.  Evaluate different methods of disposing of polymers. | 0.2 weeks | **Students should be able to:**   * explain why polyalkenes cannot be hydrolysed and so are non-biodegradable * explain why polyesters and polyamides can be hydrolysed and so are biodegradable * evaluate the advantages and disadvantages of different methods of disposing of polymers. | * Students can create a summary table to compare and explain the biodegradability of different types of polymers (AO1 - Demonstrate knowledge and understanding of scientific ideas). * Students can research and summarise different methods of disposing of polymers, including recycling, considering advantages, disadvantages and sustainability (AO3 - Analyse, interpret and evaluate scientific information). | * January 2013 Unit 4 Question 4b, 4c and 4d (QW13.4.04) * CHM4 Specimen Paper Question 5d (QSP.4.05) * June 2002 Unit 4 Question 7 (QS02.4.07) * June 2004 Unit 4 Question 5a and 5c (QS04.4.05) * SAM A-level Paper 2 (set 1) Question 7 | *Chemistry Review* article: Reclaiming plastic waste (Volume 23, edition 2)  Video on recycling plastics: <http://www.rsc.org/learn-chemistry/resource/res00001347/recycling-plastics> |

### 3.3.13 Amino acids, proteins and DNA

Amino acids, proteins and DNA are the molecules of life. In this section, the structure and bonding in these molecules and the way they interact is studied. Drug action is also considered

Prior knowledge:

**AS Chemistry**

- 3.1.3.7 – Forces between molecules.

- 3.3.1.1 – Nomenclature.

**A-level Chemistry**

- 3.3.9 – Carboxylic acids.

- 3.3.11 – Amines.

- 3.3.16 – Chromatography (you might wish to teach this section before using it to test amino acids by thin-layer chromatography here).

3.3.13.1 Amino acids

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the structure of amino acids.  Draw the structure of given amino acids in acidic solution, alkaline solution and as zwitterions. | 0.3 weeks | **Students should be able to:**   * draw the structure of given amino acids in acidic solution, alkaline solution and as zwitterions. | * Given the structure of the amino acid, students show draw the structure of the species formed in acidic solution, alkaline solution and as a zwitterion (AO2 - Apply knowledge and understanding). | * June 2013 Unit 4 Question 6 (QS13.4.06) * January 2012 Unit 4 Question 7 (QS12.4.07) * January 2005 Unit 4 Question 2 (QW05.4.02) | Structure of amino acids (rotatable) [https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/#](https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/)  RSC resource on basic biochemistry <http://www.rsc.org/Education/Teachers/Resources/cfb/proteins.htm> |

3.3.13.2 Proteins

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the structure of proteins.  Understand how peptide links can be hydrolysed to release amino acids.  Know how to use thin-layer chromatography to separate and identify amino acids. | 0.5 weeks | **Students should be able to:**   * describe the primary, secondary and tertiary structure of proteins, including the importance of hydrogen bonds and S-S bonds * draw the structure of peptides formed from amino acids * know that peptide link can be hydrolysed producing amino acids * identify the amino acids given when a peptide is hydrolysed * know that amino acids can be separated and identified by thin-layer chromatography, including the use of Rf values. | * Draw the structure of peptides formed from joining amino acids together (AO2 - Apply knowledge and understanding). * Identify amino acids formed when peptides are hydrolysed (AO2 - Apply knowledge and understanding). * Identify the primary, secondary and tertiary parts of the structure of some proteins (AO2 - Apply knowledge and understanding). * Students can carry out some thin-layer chromatography of some amino acids to identify an unknown amino acid (AO2 - Apply knowledge and understanding; AT i - Use thin-layer or paper chromatography). | * January 2010 Unit 4 Question 6 (QW10.4.06) * SAM A level Paper 2 Questions 5 * June 2011 Unit 4 Question 4c (QS11.4.04) * January 2011 Unit 4 Question 4f (QW11.4.04) | Structure of amino acids and proteins (rotatable) [https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/#](https://undergrad-ed.chemistry.ohio-state.edu/jmol-viewer/)  RSC resource on basic biochemistry <http://www.rsc.org/Education/Teachers/Resources/cfb/proteins.htm>  AQA Biochemistry Teachers’ Notes (covers 3.3.13):  <http://www.aqa.org.uk/resources/science/as-and-a-level/chemistry-7404-7405/teach/teaching-notes> |

3.3.13.3 Enzymes

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the structure of enzymes.  Understand the action of enzymes in terms of active sites.  Understand the principle of drug action and the use of computer aided design. | 0.2 weeks | **Students should be able to:**   * explain that enzymes are proteins which act through a stereospecific active site that binds to a substrate * explain how drugs, which can be designed with the aid of computers, can act to inhibit enzymes by blocking active sites, but that the correct enantiomer is required. | * Use a right handed glove with their right/left hands to model enzyme action (AO2 - Apply knowledge and understanding). |  | RSC resource on basic biochemistry of enzymes <http://www.rsc.org/Education/Teachers/Resources/cfb/enzymes.htm>  Useful animations on action of enzymes (eg hydrolysis of sucrose) <http://doctorprodigious.wordpress.com/hd-animations/> |

3.3.13.4 DNA

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand the structure of the components of DNA (given on data sheet).  Understand the nature of nucleotides.  Understand the structure of single DNA strands and the arrangement of these together in the double helix structure. | 0.3 weeks | **Students should be able to:**   * identify the components of DNA * explain how the two DNA strands interact with hydrogen bonds between base pairs. | * Make a 2D or 3D model of DNA using cut out components (AO2 - Apply knowledge and understanding). * Label a diagram of DNA to show the components and the hydrogen bonding between base pairs (AO1 - Demonstrate knowledge and understanding of scientific ideas). | * SAM A-level Paper 2 (set 1) Question 8 | How Stuff Works on the structure of DNA <http://science.howstuffworks.com/life/cellular-microscopic/dna1.htm>  Simple animation showing the structure of DNA: <http://www.youtube.com/watch?v=qy8dk5iS1f0>  Useful animations on biochemistry <http://doctorprodigious.wordpress.com/hd-animations/>  *Chemistry review* article: Why is DNA helical? (Volume 1, edition 1) |

3.3.13.5 Action of anti-cancer drugs

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Understand how DNA replicates and how anti-cancer drug cisplatin prevents this. | 0.2 weeks | **Students should be able to:**   * describe how DNA replicates in simple terms * explain how the anti-cancer drug cisplatin prevents DNA replication * explain why some drugs can have adverse effects and appreciate the balance between benefits and adverse effects of any drug. | * Write notes to accompany a sequence of diagrams showing DNA replication (AO1 - Demonstrate knowledge and understanding of scientific ideas). * Write notes to accompany a diagram showing the action of cisplatin (AO1 - Demonstrate knowledge and understanding of scientific ideas). * Evaluate the benefits and adverse effects of using drugs such as cisplatin (AO3 - Analyse, interpret and evaluate scientific information). | * SAM A-level Paper 2 (set 1) Question 8 | Useful animations on biochemistry (DNA replication) <http://doctorprodigious.wordpress.com/hd-animations/>  Youtube video on action of cisplatin <http://www.youtube.com/watch?v=Wq_up2uQRDo>  Cisplatin – molecule of the month <http://www.chm.bris.ac.uk/motm/cisplatin/htmlonly/>  *Chemistry review* article: Metals in medicine (Volume 8, edition 2)  *Chemistry review* article: Curing cancer with chemistry (Volume 18, edition 3  *Chemistry review* article: Cisplatin: from accidental discovery to wonder drug (Volume 21, edition 4) |

### 3.3.14 Organic synthesis

The formation of new organic compounds by multi-step syntheses using reactions included in the specification is covered in this section

Prior knowledge:

**AS Chemistry**

- All organic chemistry topics.

**A-level Chemistry**

- 3.3.8–3.3.13

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Devise synthetic routes to make specified compounds. | 1.0 weeks | **Students should be able to:**   * devise synthetic routes, with up to four steps, to make specific organic compounds using the reactions in the specification * explain why processes are designed to avoid solvents, non-hazardous starting materials and have steps with high atom economy. | * Devise synthetic routes, including reaction conditions, to make organic compounds using reactions in the specification (AO2 - Apply knowledge and understanding). * Describe features of processes that improve sustainability (A03 - Analyse, interpret and evaluate scientific information). | * Specimen Paper CHM4 Question 8 (QSP.4.08) * Specimen Paper CHM4 Question 9 (QSP.4.09) * June 2006 Unit 4 Question 6 (QS06.4.06) * January 2003 Unit 4 Question 7 (QW03.4.07) * June 2002 Unit 4 Question 7 (QS02.4.07) | RSC synthesis resource  <http://www.rsc.org/learn-chemistry/resource/res00000003/synthesis-explorer>  *Chemistry review* article: New tricks for stacking bricks: modern approaches to organic synthesis (Volume 12, edition 3)  *Chemistry review* article: Salbutamol: saving your breath (Volume 18, edition 4) |

### 3.3.15 Nuclear magnetic resonance spectroscopy

Chemists use a variety of techniques to deduce the structure of compounds. In this section, nuclear magnetic resonance spectroscopy is added to mass spectrometry and infrared spectroscopy as an analytical technique. The emphasis is on the use of analytical data to solve problems rather than on spectroscopic theory.

Prior knowledge:

**AS Chemistry**

- 3.3.1.1 – Nomenclature.

- 3.3.6 – Organic analysis.

This section could be taught before the A-level organic chemistry topics allowing the technique to be re-visited and to be part of practice questions throughout the teaching of the A-level organic topics.

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Using 1H and 13C NMR to deduce information about the structure of organic molecules.  Understand similarities and differences between 1H and 13C NMR.  Understand the use of TMS and suitable solvents. | 2.0 weeks | **Students should be able to:**   * understand the use of TMS and the δ scale for chemical shift * understand the use of deuterated solvents or CCl4 * use the n+1 rule to deduce spin-spin splitting patterns of adjacent, non-equivalent protons in aliphatic compounds * deduce the structure of compounds using 1H NMR to deduce structures including the number, position, relative intensity and splitting of signals * deduce the structure of compounds using 13C NMR to deduce structures including the number and position of signals. | * Predict the number, position, relative intensity and splitting of signals in the 1H NMR spectrum of compounds (AO2 - Apply knowledge and understanding). * Predict the number and position of signals in the 13C NMR spectrum of compounds (AO2 - Apply knowledge and understanding). * Use data from NMR, and other analytical methods on the specification, to deduce the structure of compounds (AO2 - Apply knowledge and understanding; MS3.1 Translate information between graphical, numerical and algebraic forms). | * June 2013 Unit 4 Question 7 (QS13.4.07) * January 2013 Unit 4 Question 5 (QS13.4.05) * June 2012 Unit 4 Question 8 (QS12.4.08) * January 2011 Unit 4 Question 5 (QW11.4.05) * January 2003 Unit 4 Question 5 (QW03.4.05) * January 2002 Unit 4 Question 4 (QW02.4.04) | RSC Spectral School: <http://www.rsc.org/learn-chemistry/collections/spectroscopy?uol_r=3ae0be55>  RSC Spectroscopy resource: <http://www.rsc.org/learn-chemistry/resource/res00000847/spectroscopy>  Database of spectra for organic compounds <http://sdbs.db.aist.go.jp/sdbs/cgi-bin/cre_index.cgi>  CLEAPSS Spectra (Secondary Science Guide L202)  <http://www.cleapss.org.uk/secondary/secondary-science/secondary-science-guides?start=20>  (Subscription required) |

### 3.3.16 Chromatography

Chromatography provides an important method of separating and identifying components in a mixture. Different types of chromatography are used depending on the composition of mixture to be separated

Prior knowledge:

**AS level Chemistry**

- 3.3.13 Amino acids, proteins and DNA (this section requires use of thin-layer chromatography for analysis of amino acids – it could be taught before or after this section)

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| **Learning objective** | **Time taken** | **Learning outcome** | **Learning activity with opportunity to develop skills** | **Assessment opportunities** | **Resources** |
| Describe the similarities and differences between thin-layer, column and gas chromatography.  Explain how chromatography works.  Use chromatography to separate and identify substances.  **Required practical 12**  Separation of species by thin-layer chromatography | 0.6 weeks | **Students should be able to:**   * describe the similarities and differences between thin-layer, column and gas chromatography * explain how chromatography works * use retention times and Rf values to identify substances * describe the use of mass spectroscopy to analyse substances separated by gas chromatography. | * Produce a summary to compare similarities and differences between thin-layer, column and gas chromatography (AO1 - Demonstrate knowledge and understanding of scientific ideas). * Separate mixtures and identify substances (eg amino acids) by thin-layer chromatography (AO2 - Apply knowledge and understanding; AT i - Use thin-layer or paper chromatography). * Use retention time and Rf data to identify substances separated by chromatography. | * January 2011 Unit 4 Question 4f (QW11.4.04) | AQA Chromatography Teachers’ Notes:  <http://filestore.aqa.org.uk/resources/chemistry/AQA-7405-TN-CHROMATOGRAPHY.PDF>  RCS video on TLC <http://www.rsc.org/learn-chemistry/resource/res00001074/thin-layer-chromatography>  Modern Chemical Techniques RSC resource:  <http://www.rsc.org/learn-chemistry/resource/res00001301/chromatography>  *Chemistry Review* articles: How pure is your aspirin? (Volume 6, edition 3)  What is chromatography? (Volume 8, edition 2)  Antarctic atmospheric chemistry (Volume 13, edition 2)  Drugs on money (Volume 13, edition 4)  Thin-layer chromatography (Volume 14, edition 3)  Body oddities: the chemical reactions of eating (Volume 21, edition 1)  Body oddities: the chemical reactions of eating (Volume21, edition 4) |